

Page 188 5.6 How elements form compounds.

Ⓚ When 2 non-metals get together to form a molecular compound they share electrons in a covalent bond.

Ⓚ When a metal and a non-metal get together to form an ionic compound the metal loses electrons to form a positive charge and the non-metal gains electrons to form a negative charge.

Recall that Atoms become ions to become more stable.

Metals: + ion (cation) ↗ attracted to each other

Non-Metals: - ion (anion) ↘

Ex: Calcium and Fluorine. Show the Bohr diagram of the ions and how Ca gives electrons to F to form a neutral ionic compound.

$^{40}_{20}\text{Ca}$ ↓ metal

loses 2e

Ca^{2+}

← one

$^{19}_9\text{F}$ ← non-metal

gains 1e

F^{-1}

← two

CaF₂ balanced
Calcium Fluoride

Mar 26-1:39 PM