

Mixed Ionic/Covalent Compound Naming
 • Transition Metals (R.N.)

For each of the following questions, determine whether the compound is ionic or covalent and name it appropriately.

- Na_2CO_3 — Sodium carbonate — Poly ionic
- P_2O_5 — diphosphorus pentoxide
- NH_3 — nitrogen trihydride
- FeSO_4 — Iron (II) Sulfate
- SiO_2 — Silicon dioxide
- GaCl_3 — Gallium Chloride
- CoBr_2 — Cobalt (II) Bromide
- B_2H_6 — diboron tetra hydride
- CO — Carbon monoxide
- P_4 — tetraphosphide

For each of the following questions, determine whether the compound is ionic or covalent and write the appropriate formula for it.

- dinitrogen trioxide — N_2O_3
- nitrogen (dinitride) — N_2
- methane (carbon tetrahydride) — CH_4
- lithium acetate — $\text{Li}(\text{C}_2\text{H}_3\text{O}_2)$
- phosphorus trifluoride — PF_3
- vanadium (V) oxide — V_2O_5
- aluminum hydroxide — $\text{Al}(\text{OH})_3$
- zinc sulfide — ZnS
- silicon tetrafluoride — SiF_4
- silver phosphate — Ag_3PO_4

Handwritten notes and oxidation states:

- Left side: $\text{Zn}^{+2} \text{S}^{-2}$, $\text{Zn}^{+2} \text{S}^{-2}$, $\text{Zn}^{+2} \text{S}^{-2}$
- Right side: $\text{V}^{+5} \text{O}^{-2}$, $\text{V}^{+5} \text{O}^{-2}$, $\text{Al}^{+3} \text{OH}^{-1}$, $\text{Al}^{+3} \text{OH}^{-1}$

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